

PNA C–C⁺ *i*-motif: superior stability of PNA TC₈ tetraplexes compared to DNA TC₈ tetraplexes at low pH[†]

Nagendra K. Sharma and Krishna N. Ganesh*

Received (in Cambridge, UK) 17th May 2005, Accepted 20th June 2005

First published as an Advance Article on the web 13th July 2005

DOI: 10.1039/b506870c

Study of self-assembly of PNA TC₈ monitored by UV thermal transition at 295 nm indicates formation of a C–C⁺ tetraplex (*i*-motif) in acidic pH, with higher stability than the analogous dTC₈.

Telomeric DNA has guanine (G) and cytosine (C) rich DNA sequence regimes. G-rich DNA oligomers are well known to form G₄-tetrads *via* WC and HG hydrogen bond mediated cyclic structures.¹ The complementary C-rich sequences form tetramers *via* the semiprotonated C–C⁺ base pairs held by three hydrogen bonds to form parallel double strands.² Two such double strands interdigitating through C–C⁺ base pairs lead to a four-stranded *i*-motif structure (Fig. 1(a)). The opposed dipoles of exocyclic C2-carbonyl and N4-amino groups favour interaction of

consecutive base pairs by alternate stacking of the amino and carbonyl groups.³

NMR spectroscopy^{3,4} has been extensively used to characterise the solution structure of the *i*-motif in oligonucleotides d(TC₅), d(T₂C₈T₂), dA₂C₄ and dC₄A₂. Rich and co-workers⁵ have solved the *i*-motif structure in several C-rich oligonucleotides by X-ray crystallography, while Raman spectroscopy⁶ was used to characterise the *i*-motif in DNAs dTC₃ and dTC₈. In UV spectra, C and protonated C⁺ show a large absorption difference at 295 nm.⁷ Hence UV-thermal transitions monitored at 295 nm, show a reverse sigmoidal pattern, which is characteristic of C–C⁺ tetraplex formation.⁷ The thermodynamics and kinetics of *i*-motif formation in modified oligonucleotides has also been studied by UV at 295 nm.⁸

DNA and RNA have very versatile auto-association properties, the range of which extends from formation of duplexes to triplexes and tetraplexes.⁹ RNA has been shown to lack the ability to form *i*-motif structures.¹⁰ Considerable interest is now growing in the study of tetraplexing properties of mimics of natural oligonucleotides such as phosphorothioates,^{8a} LNA¹¹ and PNA.¹² While G₄ tetraplex formation was successfully demonstrated recently in PNA,^{13,14} it was reported that the PNA H–C₄A₄C₄–Lys–NH₂ did not form C–C⁺ tetraplexes at pH 7.0.¹⁴ Owing to favourable steric factors, it was shown that a PNA analogue gly-ala-PNA forms C–C⁺ complexes in a C₄-tetramer, but not in a C₈-octamer.¹⁵ Thus, *no* reports exist so far on successful C–C⁺ tetraplexing properties of unmodified *aeg*-PNA. We herein present the first observation on C–C⁺ tetraplexing properties of unmodified PNA sequences TC₄ and TC₈, analogous to the isosequential DNA, but with higher thermal stability in the acidic pH range.

To study the *i*-motif in PNAs, we synthesised PNAs TC_{*n*} corresponding to different lengths (Table 1). TC₂ (PNA 1), TC₃ (PNA 2), TC₄ (PNA 3) and TC₈ (PNA 4) were synthesized by standard procedures on solid phase method using Boc-chemistry (for details see ESI[†]). For comparative study, the DNA sequences d(TC₈) and d(TC₈) were synthesized on an *ABI*-DNA synthesizer. All sequences were purified by HPLC to homogeneity and characterized by mass spectrometry.

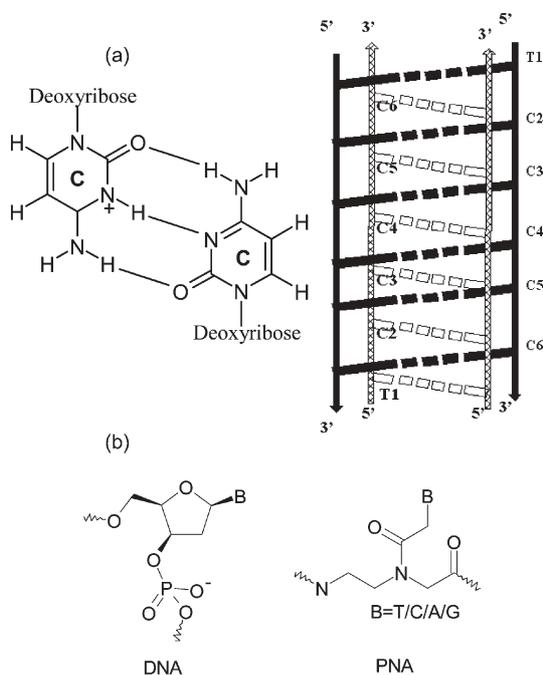


Fig. 1 (a) Schematic diagram of the *i*-motif in DNA. (b) Chemical structures of DNA and PNA.

Division of Organic Chemistry, National Chemical Laboratory, Pune, 411008, India. E-mail: kng@ems.ncl.res.in; Fax: 9120589153; Tel: 9120589153

[†] Electronic supplementary information (ESI) available: Experimental procedures, HPLC and mass spectra of PNAs 1–4, pH dependent UV spectra and UV–*T*_m measurements. See <http://dx.doi.org/10.1039/b506870c>

Table 1 Oligomers for the study of the *i*-motif of PNA

Sequences of PNA/DNA	
1	H ₂ N–T–C–C–βala–COOH
2	H ₂ N–T–C–C–βala–COOH
3	H ₂ N–T–C–C–C–βala–COOH
4	AcHN–Lys–T–C–C–C–C–C–C–CONH ₂
5	d(TCCCC)
6	d(TCCCCCCCC)

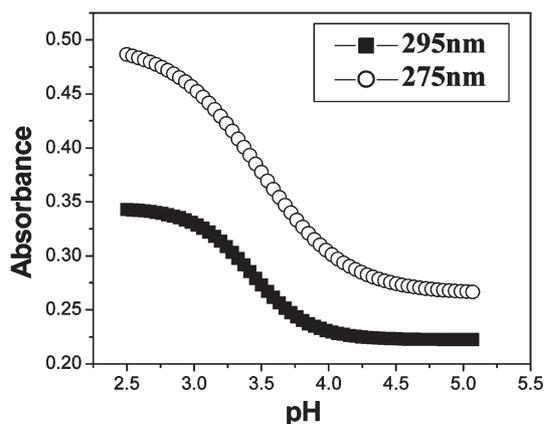


Fig. 2 UV absorbance at 275 and 295 nm of PNA 4 as a function of pH.

The UV spectra of PNA 4 were recorded at different pH values in the range 2.8–5.1 at 25 °C. The band at 275 nm found at acidic pH 2.8 slowly decreased in intensity and shifted to lower wavelength at 260 nm, with increase in pH (see ESI†). Earlier, it had been observed that the difference in absorbance spectra of protonated and non-protonated cytosine in DNA/RNA is maximum in the region 290–295 nm.¹⁰ Fig. 2 shows a plot of UV absorbance at 275 and 295 nm in PNA 4 as a function of pH and the absorbance differences between protonated and non-protonated C in PNA are greater at 275 nm. From these data, the pK_a for N3 of C in PNA is obtained as 3.45, which is significantly lower than the pK_a of 4.8 reported for N3 of C in DNA/RNA.¹⁰

The formation of C–C⁺ tetraplexes from PNAs 3, 4 and d(TC)₈ 6 at pH 3.0, 4.5, 5.0, 6.5 and 7.0 were monitored at 295 nm, for a true comparison with the tetraplex formation in d(TC)_n as per the reported procedures.^{7,10} The temperature dependent UV-absorbance results obtained are shown in Fig. 3.

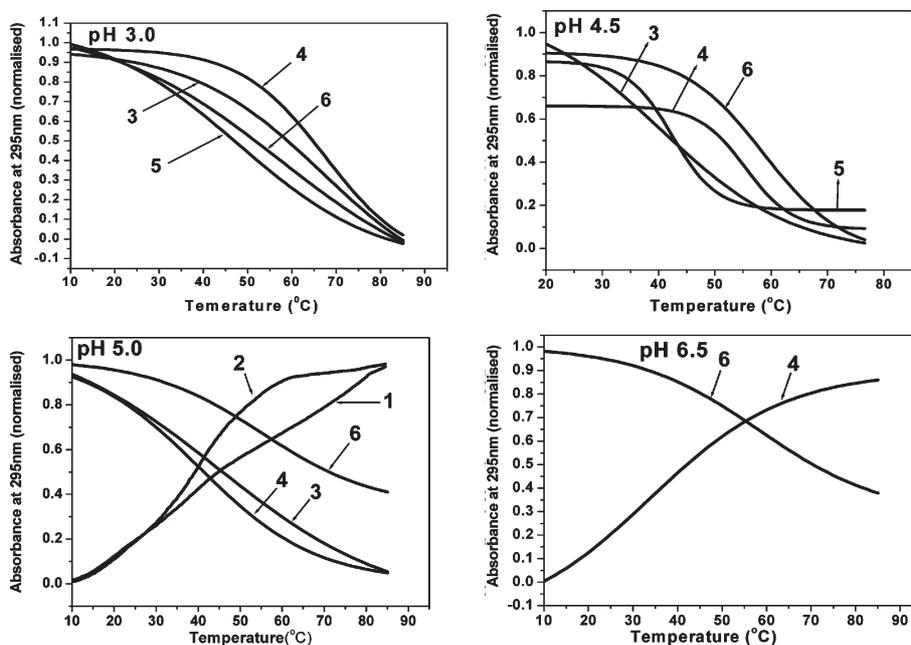


Fig. 3 UV– T_m of PNA and DNA at different pH values.

Table 2 pH Dependent T_m of TC₈ in PNA and DNA^a

pH	T_m (°C) at varying pH					
	3.0	4.5	5.0	6.0	6.5	7.0
PNA 4	67.4	55.0	46.0	nf	nf	nf
DNA 6	58.4	58.7	55.7	50.4	52.0	nf

^a nf indicates not formed.

The melting experiments were done in 100 mM sodium acetate buffer for the pH range 3.0–5.0 and 10 mM phosphate buffer for the pH range 6.0–7.0. The successful formation of tetraplexes in different sequences was indicated by observance of negative sigmoidal transitions (Fig. 3). Accurate T_m values were obtained from the first derivative curves and the T_m data for PNA 4 and DNA 6 are shown in Table 2. The PNAs 1 (TC₂) and 2 (TC₃) failed to show tetraplex formation at any of the pH conditions. PNAs 3 and 4 showed formation of strong C–C⁺ tetraplexes at pH 3 and 4.5, respectively. Significantly, these PNA C–C⁺ tetraplexes were much more stabilised (by 10–20 °C) compared to the analogous DNA C–C⁺ tetraplexes. The stability of PNA C–C⁺ tetraplexes were also dependent on pH. A comparison of pH dependent T_m of different PNA and DNA C-oligomers (Table 2) reveals that PNAs 3 and 4 form tetraplexes only in the acidic regime, up to pH 5.5. At pH 5.0, the PNA C-oligomers 3 and 4 form tetraplexes while at pH 6.0, no tetraplex formation is observed for these oligomers. This is seen from the reversal of melting curves for PNA oligomers at pH higher than 6 (Fig. 3). In comparison, the isosequential DNA C-oligomers 5 and 6 show tetraplex formation up to pH 6.5. Both PNA and DNA C-oligomers fail to form tetraplexes at pH 7.0. The pH effect on tetraplex stability is more drastic for PNA C-oligomers with $\partial T_m / \partial pH$ being 10, while that for DNA is only about 3. The difference in the cut-off pH for tetraplex formation in PNA (5.5)

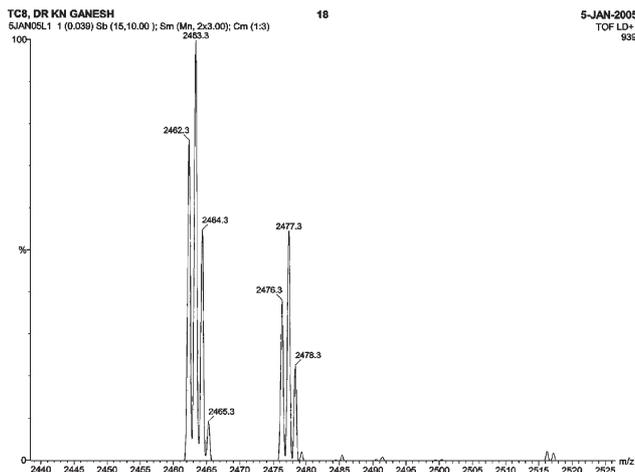


Fig. 4 MALDI-TOF mass spectrum of PNA 4.

and DNA (6.5) is perhaps a reflection of the lower pK_a of N3-C in PNA (Fig. 2) compared to that in d(TC)_n.¹⁰

The identity of PNA C-oligomers, is supported by mass spectral data (Fig. 4). The MALDI-TOF mass spectra of PNA 4 TC₈ oligomer exhibited two sets of peaks separated by 14 mass units. While the cluster at m/z 2463 corresponds to the calculated (C₉₉H₁₃₇N₄₇O₃₀) mass of M⁺, the cluster at m/z 2477 corresponds to (4M + 2H⁺ + 2Na⁺)⁴⁺.

In summary, this communication demonstrates the hitherto unknown formation of the C-C⁺ tetraplex in unmodified C-oligomeric PNAs. It is shown that in the acidic pH 3.0–5.0 range, PNA C-C⁺ tetraplexes possess significantly higher stability compared to analogous DNA C-C⁺ tetraplexes. Recently, it was reported¹⁴ that the PNA C₄A₄C₄-Lys-NH₂ did not show formation of C-C⁺ tetraplexes at pH 7.0. Up to now, no modified DNAs or their analogues have been known to form a more stable *i*-motif than natural DNA.^{2,7} In light of this and the current interest in modified peptide nucleic acid analogues,¹⁶ the first observation and characterization of C-C⁺ tetraplexes reported

here, holds promise to further examine the role of the PNA backbone structure in tetraplex formation.‡ The effect of modified backbones and sequences on influencing the self-assembling properties of nucleic acids has current importance in the development of practical applications for therapeutics and diagnostics.^{11,12}

N. K. S. thanks UGC-CSIR, New Delhi for a research fellowship. K. N. G. is an Honorary Professor of the Jawaharlal Nehru Centre for Advanced Scientific Research, Bangalore.

Notes and references

‡ Though the present results do not give a direct evidence for interdigitation of base pairs, the UV characteristics similar to that of DNA C-C⁺ tetraplexes, suggest that PNA TC₈ may have a similar arrangement. A similar structure is proposed for gly-ala PNA¹⁵

- 1 S. Neidle and G. N. Parkinson, *Curr. Opin. Struct. Biol.*, 2003, **13**, 275–283.
- 2 M. Guéron and J.-L. Leroy, *Curr. Opin. Struct. Biol.*, 2000, **10**, 326–331.
- 3 K. Gehring, J.-L. Leroy and M. Guéron, *Nature*, 1993, **363**, 561–564.
- 4 N. Esmaili and J.-L. Leroy, *Nucleic Acids Res.*, 2005, **33**, 213–224.
- 5 (a) L. Chen, L. Cai, X. Zhang and A. Rich, *Biochemistry*, 1994, **33**, 13540–13546; (b) C. H. Kang, I. Berger, C. Lockshin, R. Ratliff, R. Moyzis and A. Rich, *Proc. Natl. Acad. Sci. USA*, 1995, **92**, 3874–3878; (c) L. Cai, L. Chen, S. Raghavan, R. Ratliff, R. Moyzis and A. Rich, *Nucleic Acids Res.*, 1998, **26**, 4696–4705.
- 6 J. M. Benevides, C. H. Kang and G. J. Thomas, Jr., *Biochemistry*, 1996, **35**, 5747–5755.
- 7 A. T. Phan and J.-L. Mergny, *Nucleic Acids Res.*, 2002, **30**, 4618.
- 8 (a) J. L. Mergny and L. Lacroix, *Nucleic Acids Res.*, 1998, **26**, 4797–4803; (b) W. Lia, P. Wua, T. Ohmichia and N. Sugimotoa, *FEBS Lett.*, 2002, **526**, 77–81.
- 9 D. E. Gilbert and J. Feigon, *Curr. Opin. Struct. Biol.*, 1999, **9**, 305–314.
- 10 L. Lacroix, J. L. Mergny, J. L. Leroy and C. Helene, *Biochemistry*, 1996, **35**, 8715.
- 11 J. Wengel, *Chem. Commun.*, 2001, 1419–1924.
- 12 P. E. Nielsen and K. N. Ganesh, *Curr. Org. Chem.*, 2000, **4**, 1931–1941.
- 13 Y. Krishnan-Ghosh, E. Stephens and S. Balasubramanian, *J. Am. Chem. Soc.*, 2004, **125**, 11009–11016.
- 14 B. Datta, M. E. Bier, S. Roy and B. A. Armitage, *J. Am. Chem. Soc.*, 2005, **127**, 4199–4207.
- 15 U. Diederichsen, *Angew. Chem., Int. Ed.*, 1998, **37**, 2273–2276.
- 16 V. A. Kumar and K. N. Ganesh, *Acc. Chem. Res.*, 2005, **38**, 404–412.